

Chemistry Lab: Precipitation Reactions & Solubility

Name: _____ Period: _____ Date: _____

Data

Record precipitate formed. Describe precipitate. If there is no reaction, place a NR in the circle.

	1 calcium chloride	2 zinc acetate	3 silver nitrate	4 copper (II) sulfate
A potassium iodide	<input type="text"/>	<input type="text"/>	<input type="text"/>	<input type="text"/>
B sodium hydroxide	<input type="text"/>	<input type="text"/>	<input type="text"/>	<input type="text"/>
C sodium carbonate	<input type="text"/>	<input type="text"/>	<input type="text"/>	<input type="text"/>
D sodium phosphate	<input type="text"/>	<input type="text"/>	<input type="text"/>	<input type="text"/>

Analysis

Using the Analysis Table, write the balanced chemical equation, the balanced complete ionic equation, and the balanced net ionic equation for each reaction. Be sure to write the chemical formulas for each reactant and product correctly. If there was no reaction, no precipitate produced, write the chemical equation for the reactants and write NR for the products.

Post Lab Questions

1. Predict whether a reaction between magnesium iodide and sodium hydroxide will occur. If so, identify the precipitate. If not, explain why not using the solubility chart.
2. When solutions of ammonium hydroxide and potassium carbonate are mixed in a beaker, no visible evidence of a reaction are observed. Give a detailed explanation of what is going on in the beaker where the solutions were mixed.