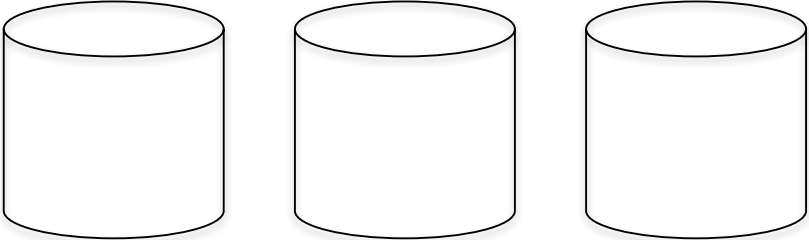


Name: _____
Period: _____ Date: _____

States of Matter & The KMT: Gases

Main Idea	Notes
<p>The Kinetic-Molecular Theory of Matter</p> <p>Kinetic: pertaining to _____</p>	<p><i>The Kinetic-Molecular Theory of Matter</i>- based on the idea that _____</p> <p>_____</p> <p>Draw & Label: Particle Drawings of Gases, Liquids, & Solids</p> <div style="display: flex; justify-content: space-around; align-items: center;"></div>
<p>The Kinetic-Molecular Theory of Gases</p> <p>move slower when it's _____!</p> <p>move faster when it's _____!</p>	<p>Five Postulates of the Kinetic-Molecular Theory of Gases: (Postulates = _____)</p> <p>1. Particles are _____</p> <p>Result: _____</p> <p>2. Collisions between gas particles, _____ are _____</p> <p>Elastic Collisions: _____</p> <p>Result: _____</p> <p>As the number of collisions increases, _____</p> <p>What causes an increase in collisions? _____</p> <p>3. The temperature of a gas, in Kelvin, is directly proportional to the _____</p> <p>If temperature doubles, _____</p> <p>At the same temperature, lighter gas particles will _____</p> <p>4. There are no _____</p> <p>5. The volume of individual gas particles is _____</p> <p>Result: _____</p>

States of Matter & The KMT: Gases

Main Idea	Notes
<i>Ideal Gases</i>	<p>The Kinetic-Molecular Theory of Gases applies to _____. Some gases nearly behave as ideal gases. But, in reality there is _____.</p> <p>Gases deviate the most from the ideal gas when _____</p> <p>Why do you think gases do not behave like ideal gases when the pressure is too high or the temperature is too low? _____</p>
<i>Properties of Gases Reflected by the KMT</i>	<p>Expansion: Gases do not have a _____ or a _____. They completely fill the container in which they are enclosed.</p> <p>Gases move _____ without any significant _____.</p> <p>Fluidity: Forces between gas particles are so insignificant that they _____.</p> <p>Gases, as well as liquids, _____ and are both _____.</p> <p>Density: Since gases are spaced so far apart, the density of gases is about _____ of the same gas in a liquid or solid state.</p> <p>Compressibility: Gases are _____ compressed because _____.</p> <p>Diffusion & Effusion: Gases spread out and _____ spontaneously, even without _____.</p> <p>Diffusion: the spontaneous mixture of the gas particles of two or more substances caused by _____.</p> <p>Effusion: The process of gas particles _____</p> <p>Effusion of gases is directly related to _____</p> <p>Lighter gas particles _____ than heavier gas particles.</p> <p>Diffusion & Effusion Drawings: Draw and label two diagrams that demonstrates the differences.</p>