



The Mole &
Mole Conversions

Name: _____

Period: _____ Date: _____

What is the molar mass for the following:

- a. Ag
- b. Fe_2O_3
- c. $\text{Al}_2(\text{SO}_4)_3$

How many formula units are present in 24.65 g CaCl_2 ?

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KEY

What is the molar mass for the following:

- a. Ag 107.87 g/mol
b. Fe₂O₃ (2 x 55.85) + (3 x 16.00) = 159.70 g/mol
c. Al₂(SO₄)₃ (2 x 26.98) + (3 x 32.07) + (12 x 16.00) = 342.17 g/mol

How many formula units are present in 24.65 g CaCl₂?

$$\frac{24.65 \text{ g CaCl}_2}{1} \times \frac{1 \text{ mol}}{110.98 \text{ g}} \times \frac{6.022 \times 10^{23} \text{ formula units}}{1 \text{ mol}} = 1.338 \text{ formula units of CaCl}_2$$